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Examples 10.2

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Law and Dalton's Law
of Partial Pressures

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Answers

The pressure of a gas
is directly proportional

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to the temperature if
the volume remains
constant What is the
formula for the
combined gas law?

$$P_1V_1/T_1 = P_2V_2/T_2$$

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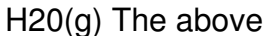
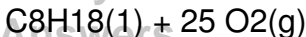
26. Ideal Gas Law
and Stoichiometry

Use the following
reaction to answer the

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next few questions: 2



reaction is the

reaction between

gasoline (octane) and

oxygen that occurs

inside automobile

engines. 29) If 4.00

moles of gasoline are

burned, what.

Gas Laws STUDY

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The volume of a gas increases as the pressure on that gas decreases. When a system that's in equilibrium is disturbed, the system adjusts itself to reduce the change. The effusion rate of a gas is...

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and its principles are

governed by the laws

of physics not the

(Answers to questions

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gases at low pressure, high temperatures, nonpolar

atoms/molecules, smaller molecules.

Dalton's Law of Partial Pressure. The

pressure of a mixture of gases is equal to

the sum of the

pressures of the

individual gases in the

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Answer Key of a gas
is directly proportional
to its temperature.

absolute zero. when
all molecular

movement stops

at -273.15 degrees C.

if the temperature and
the number of

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particles of a gas in a cylinder do not change, and the volume of the cylinder is reduced by half, the

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Chemistry

The average kinetic
energy of a gas is
expressed by the
formula: $KE_{ave} =$
 $3RT/2$ where KE_{ave}

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$\frac{3}{2}kT$ = average kinetic energy
 R = ideal gas constant
 T = absolute temperature
The average velocity or root mean square velocity of individual gas particles can be found using the formula $v_{rms} = \sqrt{\frac{3RT}{M}}$ where v_{rms} = average or root mean square velocity
 R = ideal gas

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constant T = absolute
temperature M =
molar mass

Chemistry Study
Guide for Gases -
ThoughtCo

The study guide is
divided into two
sections: vocabulary
and short answer
questions. The
vocabulary words can
be found scattered

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throughout the different instructional worksheets from this unit. The short answer questions are conceptual and meant to see if the students are able to apply what they've learned in the unit.

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Measurement: For the
following questions,
use the following
answer choices to
indicate what each
unit of measurement
is used to measure.

A. Pressure B.

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Volume C 1. K 4.

kPa A 2. atm 5. L 3.

mL 6. °C C.

Temperature 'A 7. A

8. Gases & Gas Laws

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Questions ...

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is a state of matter with no defined shape or volume. Gases have their own unique behavior depending on a variety of variables, such as temperature, pressure, and volume. While each gas is different, all gases act in a similar matter. This study guide

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AP CHEMISTRY

NOTES 5-1 THE GAS

LAWS Dalton's Law

of Partial Pressures P

total = $P_1 + P_2 + P_3$

3. . . Where pressure

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Answers
is measured in
atmospheres (atm),
kilopascals (kPa),
millimeters of mercury
(mmHg), or torr: AP
CHEMISTRY NOTES
5-1 THE GAS LAWS

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following questions,
use the following
answer choices to
indicate what each

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unit of measurement
is used to measure.

A. Pressure B.

Volume C 1. K 4.

kPa A 2. atm 5. L 3.

mL 6. °C

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